

Math Handbook Transparency Worksheet Answers

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Answers MATH HANDBOOK TRANSPARENCY MASTER Operations with Scientific Notation Addition and Subtraction Before numbers in scientific notation can be added or subtracted, the exponents must be equal. Not equal Equal (3.4 x 102) + (4.57 x 103) = (0.34 x 103) + (4.57 x 103) The decimal is moved to the left to increase the exponent.

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MATH HANDBOOK TRANSPARENCY WORKSHEET Operations with Scientific Notation I. Perform the following operations and express the answers in scientific notation. a. (1.2 x 105) + (5.35 x 106) S S b. (6.91 X 10-2) + (2.4 x 10-3) c. (9.70 x 106) + (8.3 x 105) d. (3.67 X 102) ÷ (1.6 x 101) e. (8.41 X 10:5) ÷ (7.9 x 10-6) f. (1.33 X 105) ÷ (4.9 X 104) 2.

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Scientific notation takes the form of M x 10n where 1 < M < 10 and "n" represents the number of decimal places to be moved. Positive n Indicates the standard form Is a large number. Negative n Indicates a number between zero and one, Example 1: Convert 1,500,000 to scientific notation,

[SCIENTIFIC NOTATION Name Example 1: Convert 1,500,000 to ...](#)

Math Handbook Transparency Masters Chemistry: Matter and Change 93 Addition and Subtraction Before numbers in scientific notation can be added or subtracted, the exponents must be equal. (3.4 ! 102) * (4.57 ! 103) # (0.34 ! 103) * (4.57 ! 103) Not equal Equal The decimal is moved to the left to increase the exponent. # (0.34 * 4.57) ! 103 # 4.91 ! 103 Multiplication

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Operations With Scientific Notation Answers - Displaying top 8 worksheets found for this concept. Some of the worksheets for this concept are Operations scientific notation, Scienti c notation mixed operations es1, Writing scientific notation, Writing numbers in scientific notation, Bastien, Name date class 53 math handbook transparency master 1, Scientific notation muldiv es1, Operations in scienti c notation.

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94 Chemistry: Matter and Change Math Handbook Transparency Worksheets 1. Perform the following operations and express the answers in scientific notation. a. (1.2 105) 6(5.35 10) (1.2 105) (5.35 106) (0.12 106) (5.35 106) (0.12 5.35) 106 5.47 106 b. (6.91 10 2) (2.4 10 3) (6.91 10 2) (2.4 10 3) (6.91 10 2) (0.24 10 2)

[Scientific Notation Use with Appendix B, Scientific Notation](#)

96 Chemistry: Matter and Change Math Handbook Transparency Worksheets 1. For each of the measurements in the table below, determine if the underlined number is significant or not significant. Place a check mark in the appropriate box and in the box under the rule that you used to make your determination. 2.